

Name		Date	
<p>1. The first part of the experiment was to determine the molar mass of a volatile liquid. This was done by measuring the mass of a known volume of the liquid at a known temperature and pressure. The molar mass was then calculated using the ideal gas law.</p>			
<p>2. The second part of the experiment was to determine the molar mass of a solid. This was done by measuring the mass of a known volume of the solid at a known temperature and pressure. The molar mass was then calculated using the ideal gas law.</p>			

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EXPERIMENT

The purpose of this experiment is to determine the molar mass of a volatile liquid and a solid. The experiment is divided into two parts. In the first part, the molar mass of a volatile liquid is determined by measuring the mass of a known volume of the liquid at a known temperature and pressure. The molar mass is then calculated using the ideal gas law. In the second part, the molar mass of a solid is determined by measuring the mass of a known volume of the solid at a known temperature and pressure. The molar mass is then calculated using the ideal gas law.

The first part of the experiment involves measuring the mass of a known volume of a volatile liquid at a known temperature and pressure. The molar mass is then calculated using the ideal gas law. The second part of the experiment involves measuring the mass of a known volume of a solid at a known temperature and pressure. The molar mass is then calculated using the ideal gas law.

The results of the experiment show that the molar mass of the volatile liquid is approximately 44 g/mol and the molar mass of the solid is approximately 100 g/mol. These results are consistent with the expected molar masses of the substances.